



## Chemistry and Physics; A Manual for Students and Practitioners (Paperback)

By Joseph Struthers

Rarebooksclub.com, United States, 2012. Paperback. Book Condition: New. 246 x 189 mm. Language: English . Brand New Book \*\*\*\*\* Print on Demand \*\*\*\*\*.This historic book may have numerous typos and missing text. Purchasers can download a free scanned copy of the original book (without typos) from the publisher. Not indexed. Not illustrated. 1893 Excerpt: . alcohol is known in which the hydroxyl is joined to an unsaturated carbon atom. How are the glycols prepared? By decomposing the diatomic haloid ethers with silver oxide and water, or by making the diacetates and decomposing these with caustic alkalies:  $\text{--C, H}_4\text{C}_1, +\text{Ag}_2\text{O, H}_2\text{O} = \text{CA(OH)}_2 + 2\text{AgCl}$ :  $\text{C}_2\text{H}_4\text{O} + 2\text{NaOH} = \text{C}_2\text{H}_4(\text{OH})_2 + \text{Na}_2\text{C}_2\text{H}_3\text{O}_2$ . Caustic alkalies, acting on the haloid ethers, give unsaturated hydrocarbons. What are the properties of the glycols? They are colorless, oily liquids without odor, but with a sweetish taste. Soluble in water, alcohol, and ether, ethylene glycol is insoluble in the latter alone. Oxidized, they give first mono-and then dibasic acids. Heated with haloid acids (except  $\text{H}_2\text{SO}_4$ , they give hydrines  $(\text{C}_2\text{H}_4\text{OH})_x$ . They are like alcohol in many of their properties, but, being diatomic, they can form two classes of compounds:  $\text{--H H H--C--C--H}$ , ethylene compounds;  $x \times \text{H H...}$



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